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Concepts: Terms in this
set (57) Dalton. The

atom is a tiny,

indestructible particle

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with no internal structure. Thomson. The atom is a sphere of positive electrical charge with electrons embedded in the sphere.

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the wave and particle

models of light. You

will describe how the

frequency of light

emitted by an atom is

a unique characteristic

of that atom. You will

compare and contrast

the Bohr and quantum

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mechanical models of the atom. You will express the arrangements of electrons in atoms through orbital

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Chapter 5 - electrons in atoms (handouts)

Chapter 6 - periodic table & trends

(handouts) Chapters 7/9 - ionic bonding & naming (handouts)

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Chapters 8/9 - covalent bonding & chemical names & formulas (handouts) Chapters 8/15 - VSEPR/polar bonding/IMFs (handouts) Chapter 10 - moles (handouts) Chapter 11 - reactions (handouts) Chapter 12 ...

Science / Chapter 5 - electrons in atoms (handouts)

Chapter 5: Electrons in Atoms Models of the

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Atom Rutherford used existing ideas about the atom and proposed an atomic model in which the electrons move around the nucleus, like the planets move around the sun. Rutherford's model fails to explain why objects change color when heated.

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Jennie L. Borders.

Section 5.1 - Models of
the Atom. The

Rutherford's model of
the atom did not

explain how an atom
can emit light or the

chemical properties of
an atom. Plum Pudding

Model Rutherford's

Model. The Bohr Model.

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Cristina Chemistry I H.

Terms in this set (82)

Rutherford's model-

also called nuclear

model-did not explain

how the electrons were

arranged

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in Atoms Section 5.1 -
Models of the Atom
The Rutherford's model
of the atom did not
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Model Rutherford's
Model

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...are the way electrons
are arranged in various
orbitals around the
nuclei of atoms. Three
rules tell us how:
Aufbau principle -
electrons enter the
lowest energy first.
This causes difficulties

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because of the overlap of orbitals of different energies – follow the diagram! Pauli Exclusion Principle - at most 2 electrons per orbital ...

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How many electrons can each p orbital hold? Chapter 5: Electrons in Atoms DRAFT. 10th - 11th grade, 60 times.

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02:25. Problem 1

Objects get their colors from reflecting only certain wavelengths when hit with white light. Light reflected from a green leaf is found to have a wavelength of $4.90 \times 10^{-7} \text{ m}$. What is the frequency of the light? ...

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atoms on the periodic table tend to fill low-energy orbitals first.

Thus, many students find it confusing that, for example, the 5p orbitals fill immediately after the 4d, and immediately before the 6s. The filling order is based on observed experimental results, and has been confirmed by theoretical calculations.

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**Atomic Structure
of Atoms | CHEM
1305 Introductory ...**

Electrons are extremely small. The mass of an electron is only about $1/2000$ the mass of a proton or neutron, so electrons contribute virtually nothing to the total mass of an atom. Electrons have an ...

**4.4: The Properties
of Protons,**

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**Neutrons, and
Electrons ...**

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Learn You will compare
the wave and particle
models of light. You
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frequency of light
emitted by an atom is
a unique characteristic
of that atom. You will
compare and contrast
the Bohr and quantum
mechanical models of
the atom. You will

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express the
arrangements of
electrons in atoms
through orbital

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District**

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Although the speed of
all electromagnetic
waves in a vacuum is
the same, waves can
have different
wavelengths and

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frequencies. As you can see from the equation on the previous page, wavelength and frequency are inversely related; in other words, as one quantity increases, the other decreases.

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